

TEMPLATE_TITLE	sdf
TEMPLATE_REFERENCE	df
TEMPLATE_AUTHOR	QBMUSER

QUESTION_TYPE	READING_COMPREHENSION
QUESTION_TOPIC	New-ww_16
QUESTION_ID	Q1
CHILD_QUESTIONS	1
QUESTION_TEXT	<p>The frequency of light emitted for the transition $n = 4$ to $n = 2$ of He^+ is equal to the transition in H atom corresponding to which of the following ?</p> <p>The value of enthalpy change (ΔH) for the reaction</p> $\text{C}_2\text{H}_5\text{OH}_{(l)} + 3\text{O}_{2(g)} \rightarrow 2\text{CO}_{2(g)} + 3\text{H}_2\text{O}_{(l)}$ <p>at 27°C is $-1366.5 \text{ kJ mol}^{-1}$. The value of internal energy change for the above reaction at this temperature will be :</p>
DIFFICULTY_LEVEL	UNDEFINED

QUESTION_TYPE	MULTIPLE_CHOICE_QUESTION
QUESTION_TOPIC	New-ww_16
QUESTION_ID	Q2
PARENT_PASSAGE_ID	Q1
QUESTION_TEXT	<p>The molality of a urea solution in which 0.0100 g of urea, $[(\text{NH}_2)_2\text{CO}]$ is added to 0.3000 dm^3 of water at STP is :</p>
NUMBER_OF_OPTIONS	2
1	<p>Copper crystallises in fcc lattice with a unit cell edge of 361 pm. The radius of copper atom is :</p>

	<p>An acid HA ionises as</p> $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$ <p>The pH of 1.0 M solution is 5. Its dissociation constant would be :</p>
2	<p>A 5% solution of cane sugar (molar mass 342) is isotonic with 1% of a solution of an unknown solute. The molar mass of unknown solute in g/mol is :</p> <p>The K_{sp} for $\text{Cr}(\text{OH})_3$ is 1.6×10^{-30}. The molar solubility of this compound in water is :</p>
CORRECT_ANSWER	
DIFFICULTY_LEVEL	PARENT LEVEL